THE LAW OF MASS CONSERVATIO

In the 18th century scientist thought that when things burns a substance call "phlogiston" came out of them.

Then experiments with closed vessels where substances could be accurately weighed, help scientist such as Antoine Lavoiser understand that when things burn , Oxygen is added

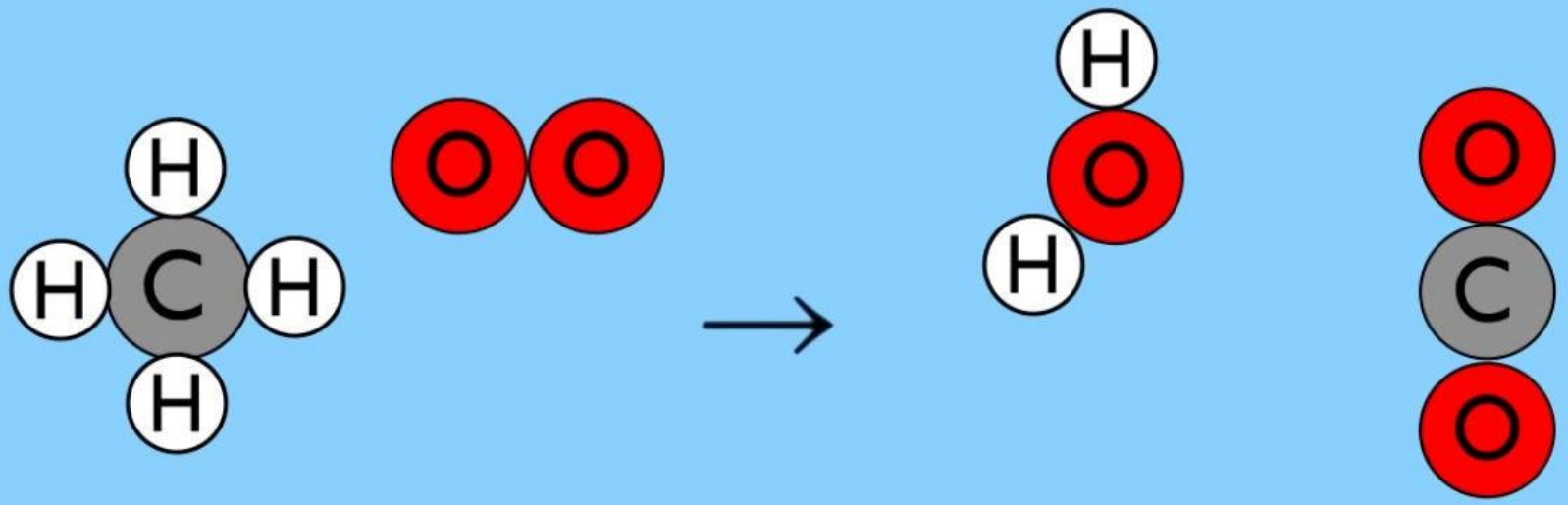
He realized that matter could be changed but not destroyed

LAW OF MASS CONSERVATION (1789)

In chemical reactions no matter is lost or grained

Reaction that we see when we using the gas cooker which uses natural gas

$$CH_4 + O_2 \rightarrow CO_2 + H_2O$$



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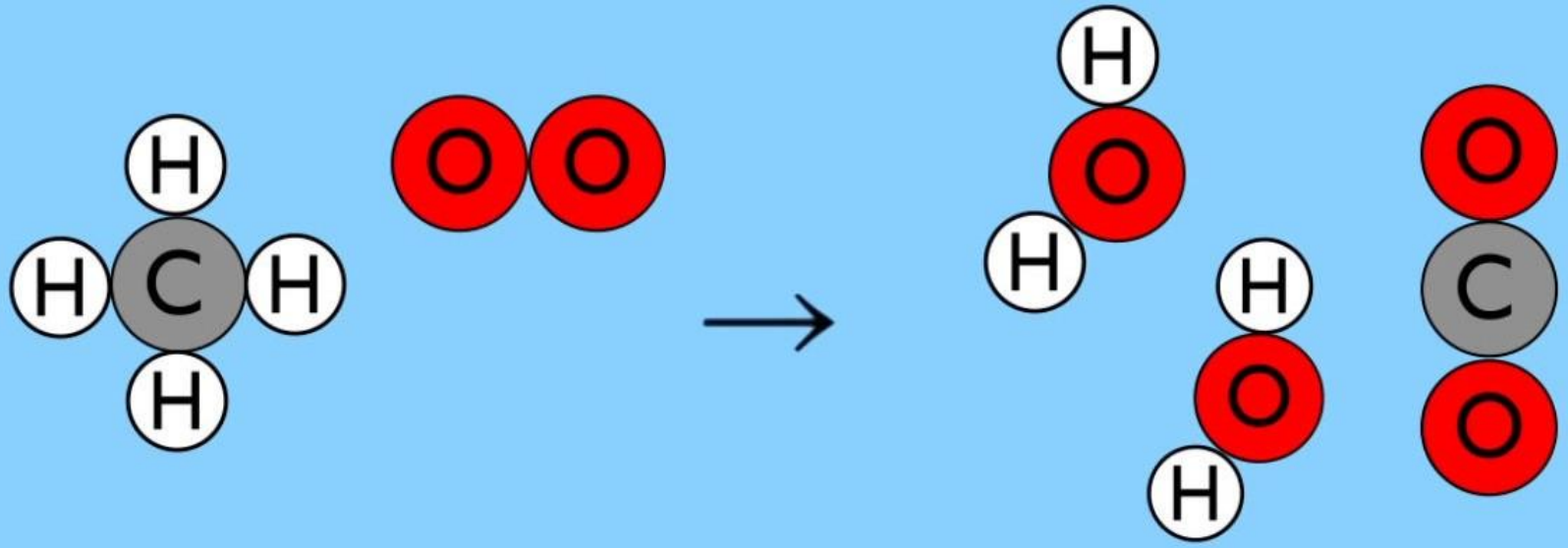
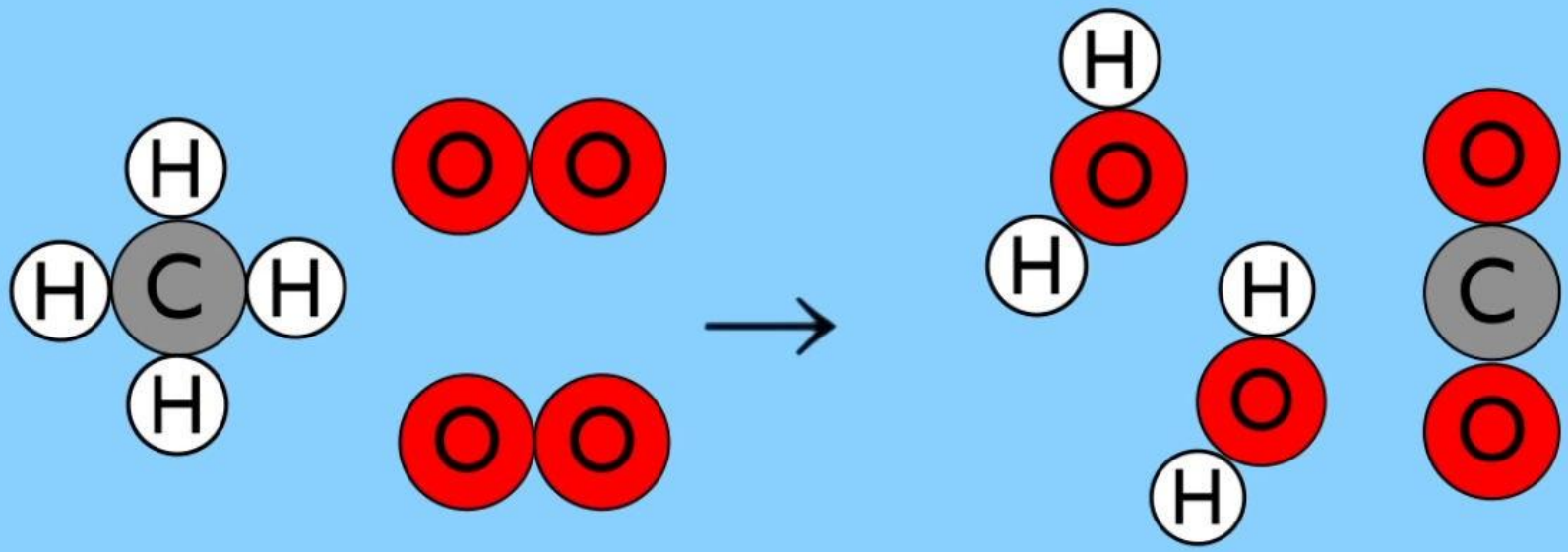


$$CH_4 + O_2 \rightarrow CO_2 + 2H_2O$$



$$CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$$

The law of conservation mass is simply saying that during chemical change there is no loss or gain of atoms

It is for this reason that
we always balance
chemical equations